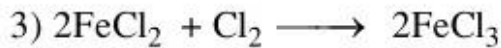
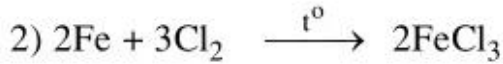
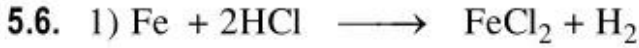


CLO



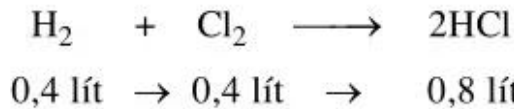
5.7. Đáp án B.

5.8. Đáp án B.

5.9. Đáp án C.

5.10. *Đáp số*:  $\approx 1,189$  tấn hay 1189 kg.

5.11. a) Trong 1 lít hỗn hợp có 0,6 lít  $\text{Cl}_2$  và 0,4 lít  $\text{H}_2$ . Clo đã lấy dư.



b) Hỗn hợp khí sau phản ứng có :  $V_{\text{HCl}} = 0,8 \text{ lít} \rightarrow 80\% \text{ thể tích}$ ,  
 $V_{\text{Cl}_2 \text{ dư}} = 0,2 \text{ lít} \rightarrow 20\% \text{ thể tích}$ .

5.12.  $n_{\text{CuCl}_2} = \frac{27}{135} = 0,2 \text{ (mol)}$

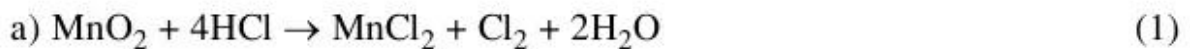


$$0,2 \text{ mol} \leftarrow 0,2 \text{ mol} \leftarrow 0,2 \text{ mol}$$

$$m_{\text{Cu}} = 64.0,2 = 12,8 \text{ (g)}$$

$$V_{\text{Cl}_2} = 22,4.0,2 = 4,48 \text{ (lít)}$$

5.13.  $n_{\text{HCl}} = \frac{7,3}{36,5} = 0,2 \text{ (mol)}$



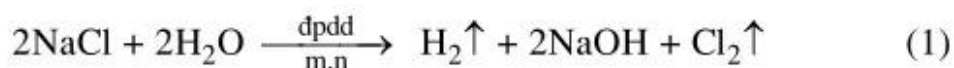
Theo (1) :  $n_{\text{Cl}_2} = \frac{0,2}{4} = 0,05 \text{ (mol)}$  ;  $V_{\text{Cl}_2} = 22,4.0,05 = 1,12 \text{ (lít)}$



$$\text{Theo (2): } n_{\text{Cl}_2} = \frac{0,2.5}{16} = 0,0625 \text{ (mol)}$$

$$V_{\text{Cl}_2} = 22,4.0,0625 = 1,4 \text{ (lít)}$$

$$\mathbf{5.14.} \quad n_{\text{Cl}_2} = \frac{560}{22,4} = 25 \text{ (mol)}$$



$$\text{Theo (1): } n_{\text{NaCl cần}} = 25.2 = 50 \text{ (mol)}$$

Khối lượng muối ăn chứa 98% NaCl cần lấy là :

$$\frac{58,5.50.100}{98} = 2984,69 \text{ (g)} \approx 2,985 \text{ kg}$$